

<b>Module code</b>	SP-2303		
<b>Module Title</b>	Quantum Mechanics and Atomic Physics		
<b>Degree/Diploma</b>	Bachelor of Science (Applied Physics)		
<b>Type of Module</b>	Major Option		
<b>Modular Credits</b>	4	<b>Total student Workload</b>	8 hours/week
		<b>Contact hours</b>	4 hours/week
<b>Prerequisite</b>	None		
<b>Anti-requisite</b>	SP-2203 Quantum Mechanics and Atomic Physics		
<b>Aims</b>			
<p>In this module, you will learn the fundamental principles of quantum mechanics and the mathematical techniques needed to state and apply them, for instance, to study the structure of atom with a single electron and many electrons. Quantum mechanics will also be used to study the bonding between atoms to form molecules.</p>			
<b>Learning Outcomes</b>			
<i>On successful completion of this module, a student will be expected to be able to:</i>			
Lower order :	30%	<ul style="list-style-type: none"> <li>- build up a sound knowledge of the mathematical ideas related to the quantum mechanical concepts</li> <li>- explain the concept of spin, spin angular momentum, its orientations, and Pauli's exclusion principle.</li> </ul>	
Middle order :	60%	<ul style="list-style-type: none"> <li>- apply quantum mechanical operators to explain the concepts of quantization of energy, barrier penetration and emission of alpha particle(s) in radioactive nuclei</li> </ul>	
Higher order:	10%	<ul style="list-style-type: none"> <li>- use quantum mechanical methods to model phenomena in physical systems including atoms, nucleus, molecules and solids.</li> </ul>	
<b>Module Contents</b>			
Quantum Mechanics:			
<ul style="list-style-type: none"> <li>- Introduction to quantum mechanics, quantum mechanical operators</li> <li>- Heisenberg uncertainty relation, Time-independent Schrödinger equation</li> <li>- Stationary states, eigenstates and eigenvalues, bound states in a potential</li> <li>- Infinite square well potential, potential step</li> <li>- Free particle, probability current, parity operator, and harmonic oscillator.</li> </ul>			
Atomic Physics:			
<ul style="list-style-type: none"> <li>- Application of Schroedinger's equation to the hydrogen atom, hydrogen spectrum</li> <li>- Origin of the quantum numbers, electronic probability density</li> <li>- Selection rules, normal Zeeman's effect, electron spin, total angular momentum</li> <li>- Exclusion principle, periodic table, electronic structure of elements, spin orbit coupling</li> <li>- X-ray spectrum and Moseley plot</li> </ul>			
<b>Assessment</b>	Formative assessment	Problem-solving, tutorials, group discussions and feedback	
	Summative assessment	Examination: 60% Coursework: 40% <ul style="list-style-type: none"> <li>- 3 assignments (20%)</li> <li>- 3 class tests (20%)</li> </ul>	