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| Module code | SP-4309 | | |
| Module Title | Introduction to Chemical Physics | | |
| Degree/Diploma | Bachelor of Science (Applied Physics) | | |
| Type of Module | Major Option | | |
| Modular Credits | 4 | Total student Workload | 12 hours/week |
| | | Contact hours | 4 hours/week |
| Prerequisite | None | | |
| Anti-requisite | None | | |
| Aims | | | |
| This module aims to introduce students to the fields of chemical physics and physical chemistry, and to demonstrate the application of physical methods and theories to chemical systems. | | | |
| Learning Outcomes | | | |
| <i>On successful completion of this module, a student will be expected to be able to:</i> | | | |
| Lower order : | 0% | | |
| Middle order : | 30% | <ul style="list-style-type: none"> - describe the main fields of study in chemical physics and physical chemistry - appreciate the relevance of physical methods and theories to chemical systems | |
| Higher order: | 70% | <ul style="list-style-type: none"> - solve chemical problems using the principles of thermodynamics and chemical kinetics - interpret the absorption and emission spectra of simple molecules - demonstrate an understanding of basic electrochemistry and electron transfer theories - choose appropriate experimental techniques for the measurement of specific physical properties - work independently and also collaboratively in a team - interpret the results of analyses, and make appropriate reports and presentations for effective communication | |
| Module Contents | | | |
| The main contents of the module are: | | | |
| <ul style="list-style-type: none"> - introductory thermodynamics and chemical kinetics (the laws of thermodynamics; kinetic theory of gases; rate equations; reaction order) - introductory molecular spectroscopy and photophysics (electronic, vibrational, and rotational spectra; fluorescence and phosphorescence; Jablonski diagrams; the Franck–Condon principle) - electrochemistry (cell thermodynamics; electrochemical kinetics; the Butler–Volmer equation; Fick’s first and second laws) - electron transfer theory (Marcus–Gerischer theory; Marcus–Hush theory) - selected experimental techniques (UV-Vis-NIR; transient absorption; voltammetry; impedance spectroscopy) | | | |
| Assessment | Formative assessment | In-class questions and homework problems will be used to test and give feedback on learning. | |
| | Summative assessment | Examination: 0% Coursework: 100% <ul style="list-style-type: none"> - Two (2) class tests (50%) - Three (3) written assignments (30%) - One (1) literature survey (10%) - One (1) oral presentation (10%) | |